

**LABORATORY REPORT SHEET**

Name \_\_\_\_\_ Date \_\_\_\_\_ Section \_\_\_\_\_

Complete the following work sheets.

In the "Observations" section, describe, in your own words, what you saw (heard, smelled, etc.) when the reagents were mixed.

In the "Equation" section, write the net ionic equation for the reaction which occurred. If you think that no reaction occurred, just write N.R. (for "No Reaction").

In the "Comments" section, explain how your observations are consistent with your net ionic equation. For example, if your equation says that  $\text{Cu}_{(s)}$  is formed in a reaction, you could comment that you observed a brown solid in the test tube; if the equation says the ammonia was released, comment on whether you could smell it, etc.

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**Reaction 1**                       $\text{H}_2\text{SO}_4 + \text{H}_2\text{O}$

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 2**                       $\text{NH}_4\text{Cl} + \text{H}_2\text{O}$

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

Experiment 10-003c  
Some Chemical Reactions

**Reaction 3**                      KOH + HCl  
Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 4**                      Pb(NO<sub>3</sub>)<sub>2</sub> + HCl  
Observations:

Net Ionic  
Equation: \_\_\_\_\_

Comments:

Observations:                      AgNO<sub>3</sub> + HCl

Net Ionic  
Equation: \_\_\_\_\_

Comments:

**Reaction 5a**

various metal nitrates + KOH

Observations:

Net Ionic Equations

Comments:

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**Reaction 5b**

products from 5a + HCl

Observations:

Net Ionic Equations

Comments:

Experiment 10-003c  
Some Chemical Reactions

**Reaction 6**       $\text{NaNO}_3 + \text{Ba}(\text{NO}_3)_2$

Observations:

Net Ionic  
Equation: \_\_\_\_\_

Comments:

$\text{K}_2\text{SO}_4 + \text{Ba}(\text{NO}_3)_2$

Observations:

Net Ionic  
Equation: \_\_\_\_\_

Comments:

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**Reaction 7a**       $\text{NaHCO}_3 + \text{HCl}$

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

**Reaction 7b**

chalk + HCl

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments

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**Reaction 8**

$\text{NH}_4\text{Cl} + \text{KOH}$

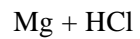
Observations:

Net Ionic Equation: \_\_\_\_\_

Comments: (be sure to account for any pH changes):

Experiment 10-003c  
Some Chemical Reactions

**Reaction 9a**



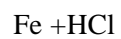
Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 9b.**

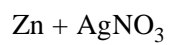


Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

**Reaction 10**



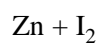
Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 11**



Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

Experiment 10-003c  
Some Chemical Reactions

**Reaction 12**                       $\text{H}_2\text{O}_2 + \text{KMnO}_4$  (in acid)

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 13**                      US penny +  $\text{HCl}_{(\text{aq})}$

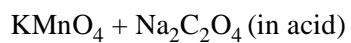
Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:



**Reaction 14**



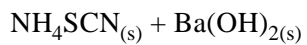
Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 15**



Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

Experiment 10-003c  
Some Chemical Reactions

**Reaction 16a**                      Mg + ??

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments:

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**Reaction 16b**                      product 16a + H<sub>2</sub>O

Observations:

Net Ionic Equation: \_\_\_\_\_

Comments: