

Results and Calculations:

Draw standard curve and use Beer's law to determine concentration of ASA in your aspirin sample.

For Standard Curve: x-axis \Rightarrow concentration ASA (M)
y-axis \Rightarrow absorbance (no units) at λ_{max} nm

e.g. A 0.0252 g sample of ASA was dissolved in base and brought to volume in a 25.0 mL Concentration of ASA:

$$\frac{0.0252 \text{ g ASA}}{180.16 \frac{\text{g}}{\text{mole}}} = 1.399 \times 10^{-4} \text{ mole ASA}$$
$$\frac{1.399 \times 10^{-4} \text{ mole ASA}}{0.0250 \text{ L}} = 0.005595 \text{ M}$$

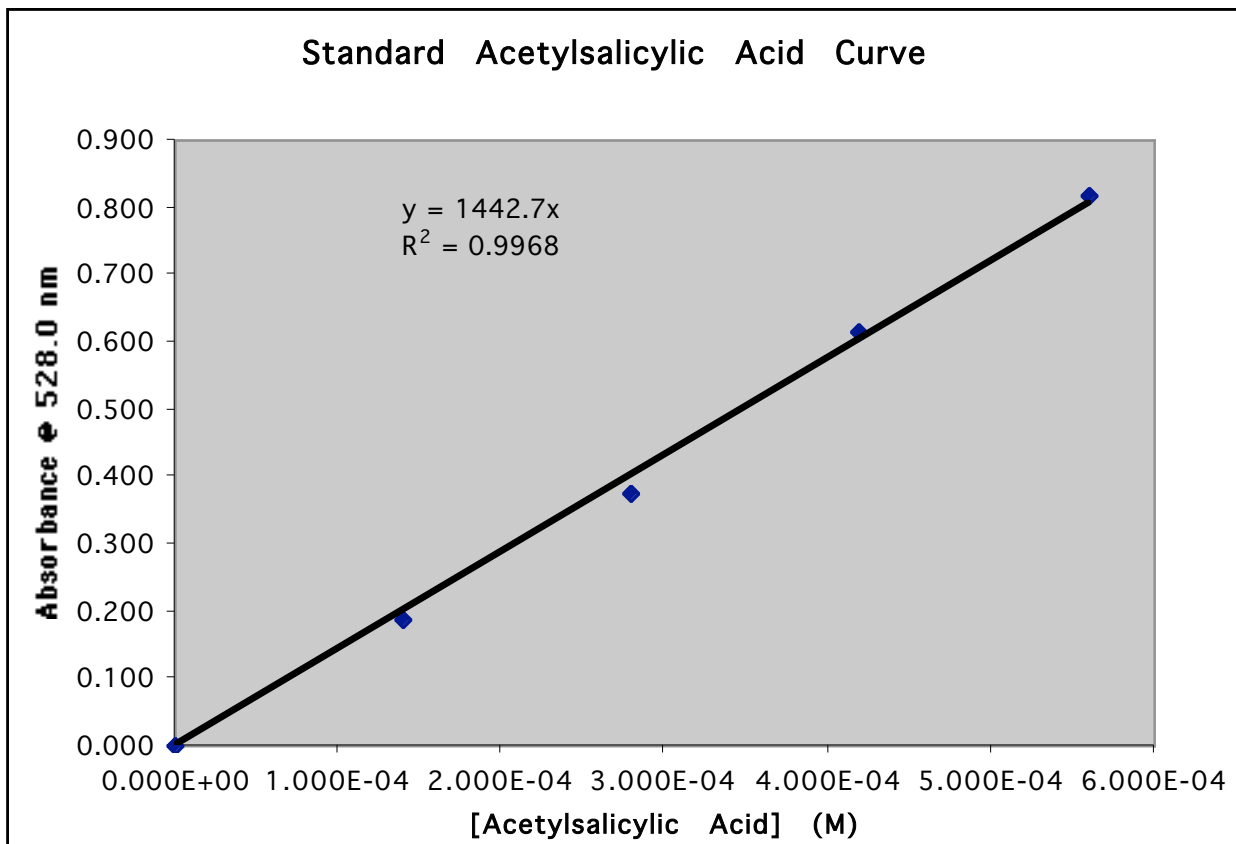
For standard series:

$$(0.005595 \text{ M})(0.00025 \text{ L}) = (x)(0.0100 \text{ L})$$
$$x = 1.399 \times 10^{-4} \text{ M}$$

Table 1: Standard ASA curve data.

Volume ASA (mL)	Concentration ASA (M)	Absorbance @ 528.0 nm
0.00	0.000	0.000
0.25	1.399×10^{-4}	0.186
0.50	2.798×10^{-4}	0.374
0.75	4.196×10^{-4}	0.616
1.00	5.595×10^{-4}	0.818

Figure 1: Standard Acetylsalicylic Acid Curve



Use standard curve to determine ϵ and use Beer's law to solve for concentration of ASA in your sample.

From Standard Curve:

$$\epsilon = 1442.7 \text{ cm}^{-1}\text{M}^{-1}$$

Beer's law: $A = \epsilon lc \Rightarrow c = \frac{A}{\epsilon l}$, $l = 1.00 \text{ cm}$ for this experiment

Table 2: Absorbance Data for Bayer aspirin tablet:

Absorbance of Aspirin sample @ 528.0 nm	Concentration of aspirin sample in cuvette (M)
0.552	3.823×10^{-4}
0.533	3.691×10^{-4}
0.530	3.670×10^{-4}
0.540	3.743×10^{-4}

Mean Concentration of Aspirin sample: $3.732 \times 10^{-4} \text{ M}$

Standard Deviation: $6.777 \times 10^{-6} \text{ M}$ (see p. 21 of lab manual)

Calculation of concentration in your aspirin sample: **Remember Dilutions**

1. Diluted 0.50 mL to 10.0 mL $\Rightarrow \frac{10.0}{0.50} = 20$

so $(3.732 \times 10^{-4} \text{ M})(20) = 7.464 \times 10^{-3} \text{ M}$

2. Total moles in 250 mL solution:

$$(7.464 \times 10^{-3} \text{ mole/L})(0.250 \text{ L}) = 1.866 \times 10^{-3} \text{ mole}$$

3. Total grams (g) of acetylsalicylic acid in sample:

$$(1.866 \times 10^{-3} \text{ mole})(180.16 \text{ g/mole}) = 0.3362 \text{ g ASA}$$

4. Total mg ASA:

$$(0.3362 \text{ g})(1000 \text{ mg/g}) = 336.2 \text{ mg ASA}$$